UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
General Certificate of Education Ordinary Level

CANDIDATE NAME

CENTRE NUMBER

CANDIDATE NUMBER

0707276106

COMBINED SCIENCE **5129/22**

Paper 2 **October/November 2012**

2 hours 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.
Write in dark blue or black pen.
You may use a soft pencil for any diagrams, graphs or rough working.
Do not use staples, paper clips, highlighters, glue or correction fluid.
DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.
A copy of the Periodic Table is printed on page 28.

At the end of the examination, fasten all your work securely together.
The number of marks is given in brackets [ ] at the end of each question or part question.

For Examiner's Use

This document consists of **24** printed pages.

For Examiner's Use

1 Use words from the list to complete the sentences below.

blood	**glands**	**kidney**	**liver**
nerves	**main**	**system**	**target**

Each word may be used once, more than once or not at all.

Hormones are chemicals that are produced by ..... .

Hormones are transported round the body by ..... .

Each hormone affects the activity of a part of the body which is called the

..... organ.

Hormones are destroyed by the ..... . [4]

2 **(a)** A weight of 2.5 N falls vertically through a distance of 2.4 m.

Calculate the work done on the weight by the force of gravity.

work done = ... unit ..... [3]

(b) The falling weight is used to rotate a coil in a magnetic field.

An e.m.f. is induced across the ends of the coil.

State two factors that affect the magnitude of the induced e.m.f.

1. .....

2. .....

[2]

3 Fig. 3.1 shows the path of a ball thrown from the top of a building.

For Examiner's Use

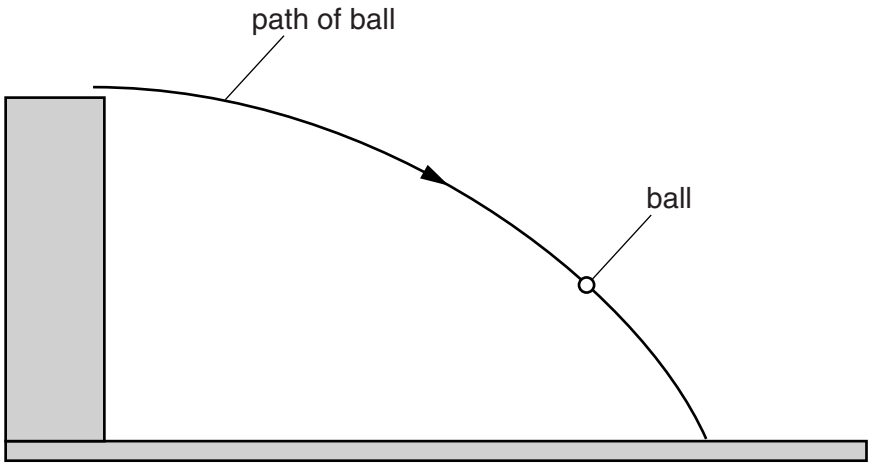


Fig. 3.1

The ball is shown at one position in its path.

(a) On Fig. 3.1, draw an arrow to show the direction of the force of gravity acting on the ball. [1]

(b) On the path of the ball shown in Fig. 3.1

(i) mark where the ball has maximum potential energy and label this point **P**, [1]

(ii) mark where the ball has maximum kinetic energy and label this point **K**. [1]

(c) The ball accelerates because of the force of gravity.

Explain what is meant by *acceleration*.

.....

..... [1]

[Turn over

4 Fig. 4.1 shows the electronic structure of a magnesium atom.

For Examiner's Use

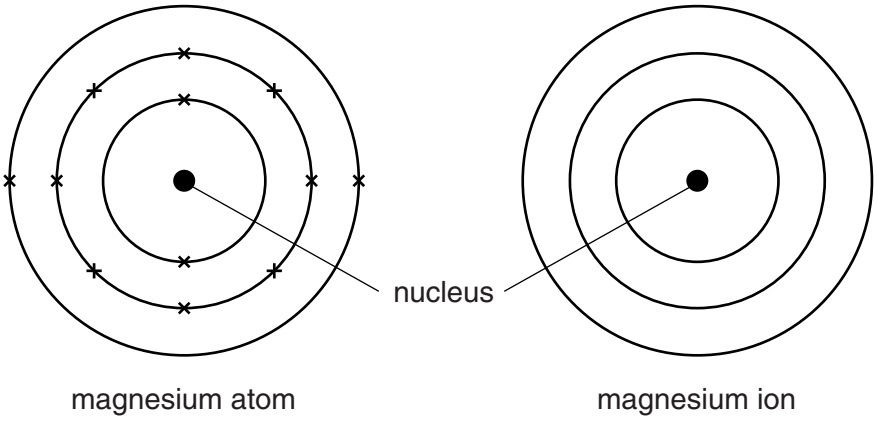


Fig. 4.1

(a) On Fig. 4.1, complete the electronic structure for the magnesium ion. [1]

(b) Magnesium burns in carbon dioxide to form magnesium oxide and carbon.

The equation for the reaction is

$$2Mg + CO_2 \longrightarrow 2MgO + C$$

The relative molecular mass, M_r, of carbon dioxide is 44.
[A_r: Mg, 24; O, 16; C, 12]

Complete the following sentences.

44 g of carbon dioxide produces .....g of magnesium oxide and .....g of carbon.

4.4 g of carbon dioxide produces .....g of magnesium oxide and .....g of carbon.

1.1 g of carbon dioxide produces .....g of magnesium oxide. [4]

(c) Magnesium oxide is a white solid with a high melting point.

State the type of bonding present in magnesium oxide.

..... [1]

5 A path is made by laying concrete slabs side-by-side.

Small gaps are left between the slabs.

The gaps are filled with sand.

Fig. 5.1 shows the slabs on a cold day.

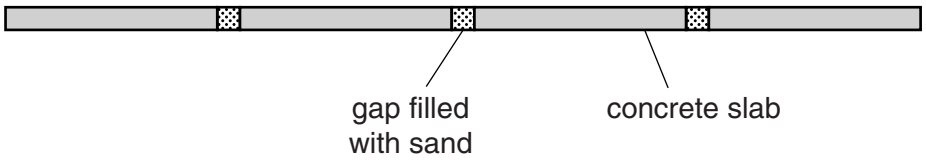


Fig. 5.1

(a) On a hot day, the gaps are smaller than on a cold day.

Explain why.

.....

..... [1]

(b) Another path is laid on a cold day with no gaps between the concrete slabs.

Suggest what may happen to this path on a very hot day.

.....

..... [1]

For Examiner's Use

6 Fig. 6.1 shows an animal cell as seen using a microscope.

For Examiner's Use

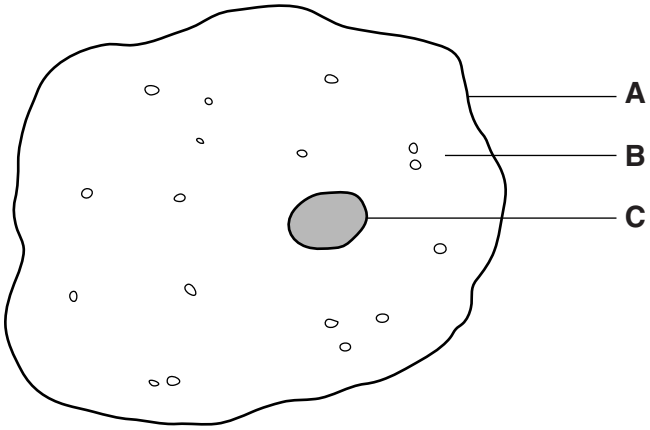


Fig. 6.1

(a) (i) In Table 6.1, name each of the labelled parts.

Table 6.1

letter	name
A	
B	
C	

[3]

(ii) State a function of part **A**.

.....

..... [1]

(b) The structure of a red blood cell is different from that of the cell shown in Fig. 6.1.

For Examiner's Use

State **two** ways in which the structure is different.

Explain how each difference helps the red blood cell to carry out its function.

difference 1 .....

.....

explanation 1 .....

.....

.....

.....

difference 2 .....

.....

explanation 2 .....

.....

.....

.....

[6]

For Examiner's Use

7 Alpha-particles, beta-particles and gamma-rays are types of emission from radioactive sources.

(a) State the type of emission that

(i) is the most penetrating, ... [1]

(ii) consists of two protons and two neutrons. ..... [1]

(b) A nucleus emits a beta-particle.

State the change that occurs in the nucleus.

..... [1]

(c) A radioactive source is used in a laboratory experiment.

State two precautions that are taken to use the source safely.

1. .....

2. .....
[2]

8 Permanent magnets and electromagnets may be used to separate magnetic materials from non-magnetic materials.

(a) State a difference between magnetic materials and non-magnetic materials.

.....

..... [1]

(b) An electromagnet is used in a simple lock.

Fig. 8.1 shows part of this lock.

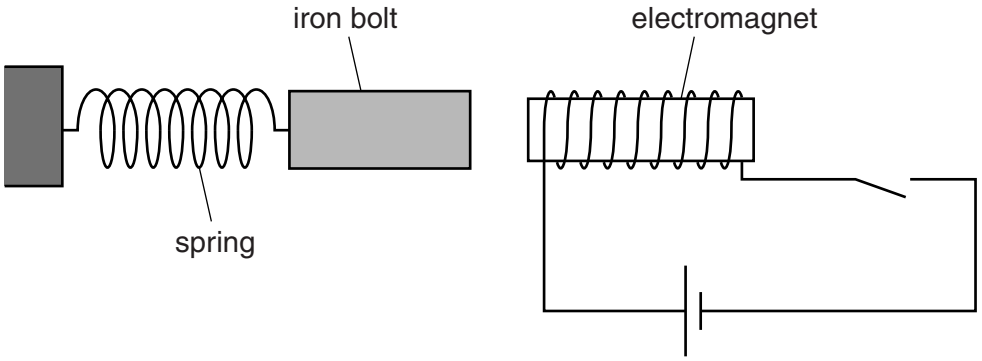


Fig. 8.1

When the current is switched on, the iron bolt is pulled towards the electromagnet to lock the door.

When the current is switched off, the spring pulls the iron bolt away from the electromagnet, unlocking the door.

(i) Suggest why the bolt is made of iron rather than steel.

.....

.....

..... [1]

(ii) The connections to the cell in Fig. 8.1 are reversed.

State the difference, if any, that this makes to the working of the lock.

.....

..... [1]

(iii) State **two** ways in which the strength of the electromagnet may be increased.

1. .....

2. .....
[2]

(c) Fig. 8.2 shows how the extension of the spring varies with the load on the spring.

For Examiner's Use

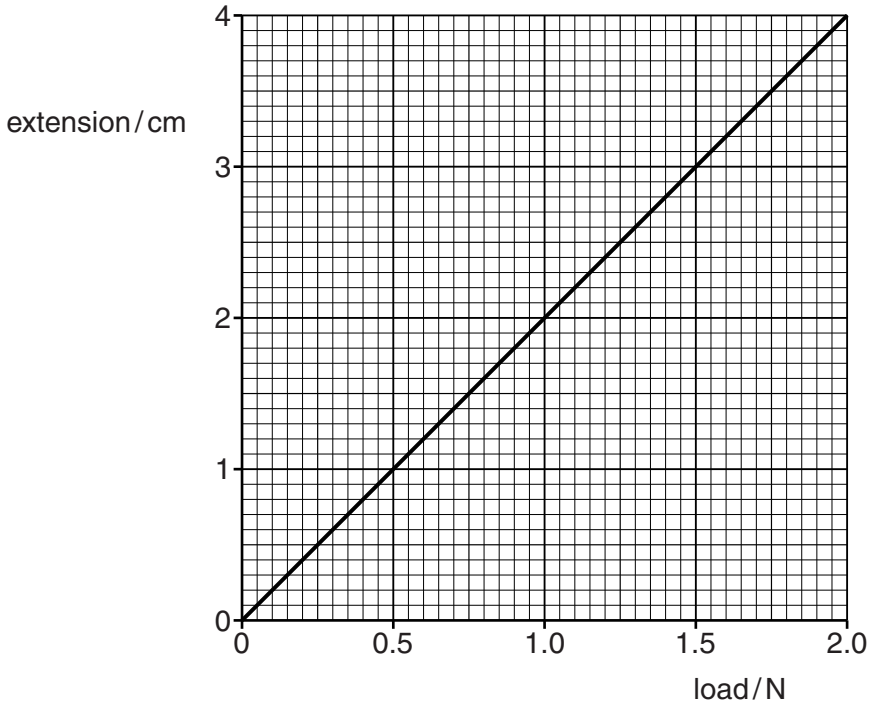


Fig. 8.2

Use Fig. 8.2 to find the load on the spring when it has an extension of 1.6 cm.

load = ... N [1]

9 Study the reaction scheme shown in Fig. 9.1.

For Examiner's Use

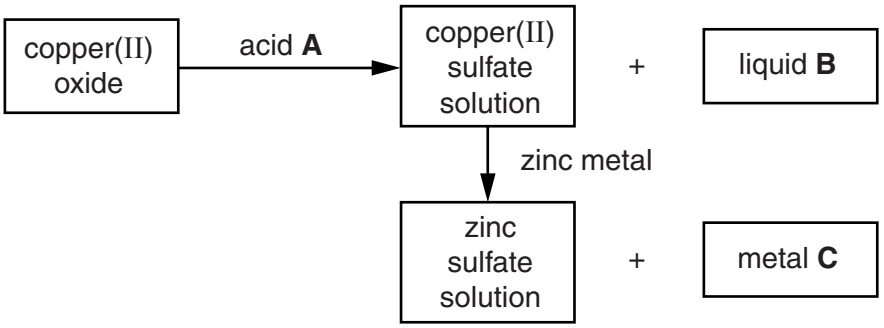


Fig. 9.1

(a) Identify **A**, **B** and **C**.

acid **A** .....

liquid **B** .....

metal **C** ..... [3]

(b) Describe how copper(II) sulfate crystals may be obtained from the copper(II) sulfate solution.

.....

.....

..... [2]

(c) State **two** general physical properties of substance **C** that show it is a metal.

1. .....

2. .....

[2]

10 In an experiment, 20 seeds of the same species are placed in each of four tubes as shown in Fig. 10.1.

For Examiner's Use

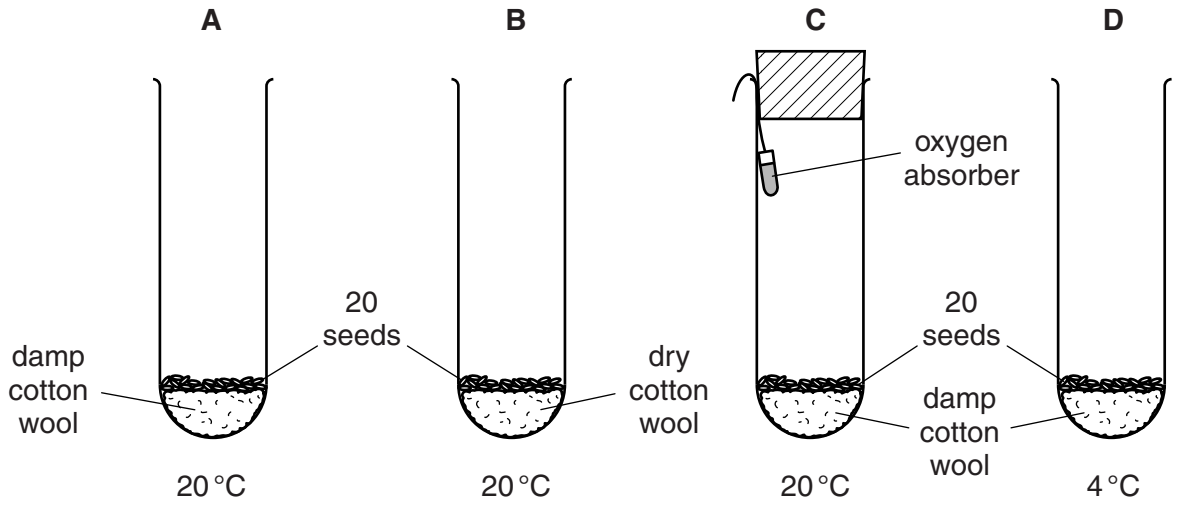


Fig. 10.1

The conditions for each set of seeds are shown in Fig. 10.1.

(a) Suggest why more than one seed is used in each tube.

.....

..... [1]

(b) After several days, all the seeds in tube **A** germinate.

None of the seeds in tubes **B**, **C** or **D** germinate.

(i) Suggest a change that could be made to tube **B** so that the seeds germinate.

.....

..... [1]

(ii) State a reason why germination does not occur in tube **C**.

Explain your answer.

reason .....

explanation .....

.....

[2]

(iii) State a reason why germination does not occur in tube **D**.

For Examiner's Use

Explain your answer.

reason .....

explanation .....

.....

[2]

11 Fig. 11.1 shows a series circuit.

For Examiner's Use

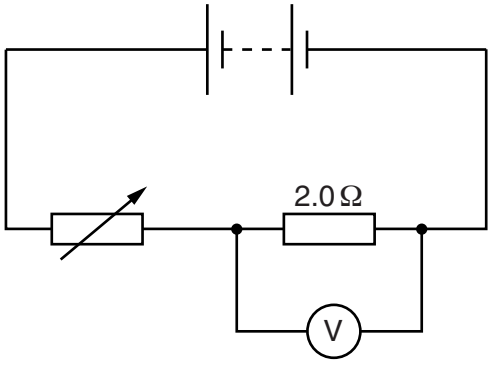


Fig. 11.1

A voltmeter measures the potential difference (p.d.) across the 2.0 Ω resistor.

(a) The variable resistor is adjusted so that the voltmeter reads 1.0 V.

(i) Calculate the current in the 2.0 Ω resistor.

current = ... A [2]

(ii) The p.d. across the battery terminals is 5.0 V.

The voltmeter reads 1.0 V.

Calculate the p.d. across the variable resistor.

p.d. = ... V [1]

(b) The resistance of the variable resistor is increased.

State what happens, if anything, to

(i) the current in the variable resistor, ... [1]

(ii) the p.d. across the 2.0 Ω resistor. ..... [1]

12 Fig. 12.1 shows the apparatus used to pass steam over heated zinc.

For Examiner's Use

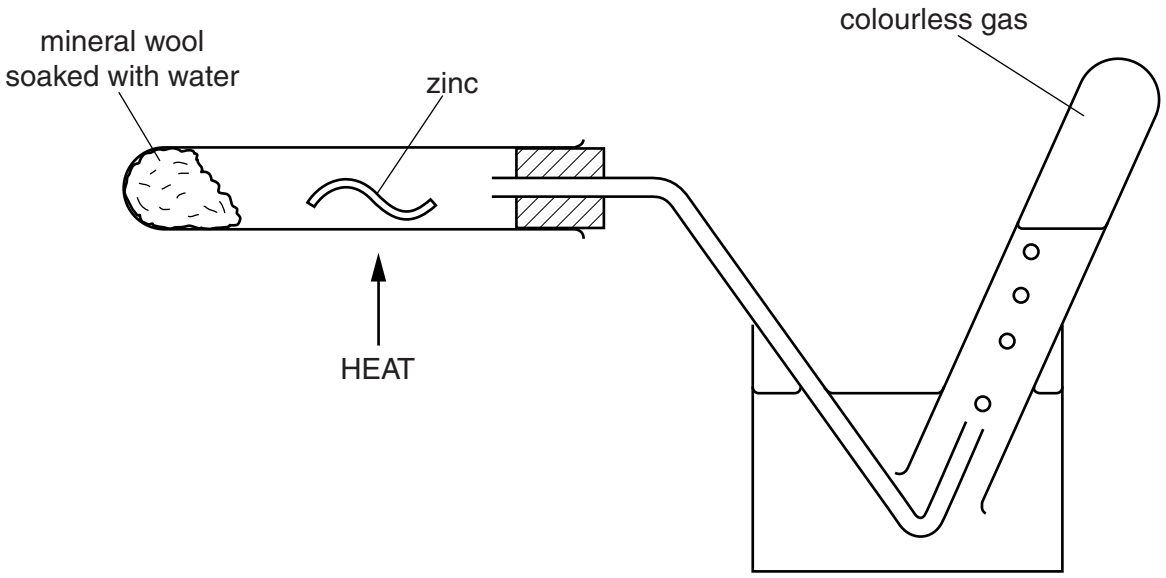


Fig. 12.1

The products of the reaction are zinc oxide and a colourless gas.

(a) Complete the equation for the reaction.

$$Zn + H_2O \longrightarrow ZnO + ...$$ [1]

(b) Explain how the equation in **(a)** shows that zinc is oxidised and steam is reduced during the reaction.

.....

.....

..... [2]

(c) Zinc is used to prevent iron from rusting.

(i) State the name of each of the two substances in air which cause iron to rust.

..... and ..... [2]

(ii) State the name of the process where iron is treated with zinc to prevent rusting.

..... [1]

[Turn over

13 Five similar fields are used for growing maize. They are treated with different quantities of nitrogen-containing fertiliser.

For Examiner's Use

The quantities of fertiliser added and the crop yields are shown in Fig. 13.1.

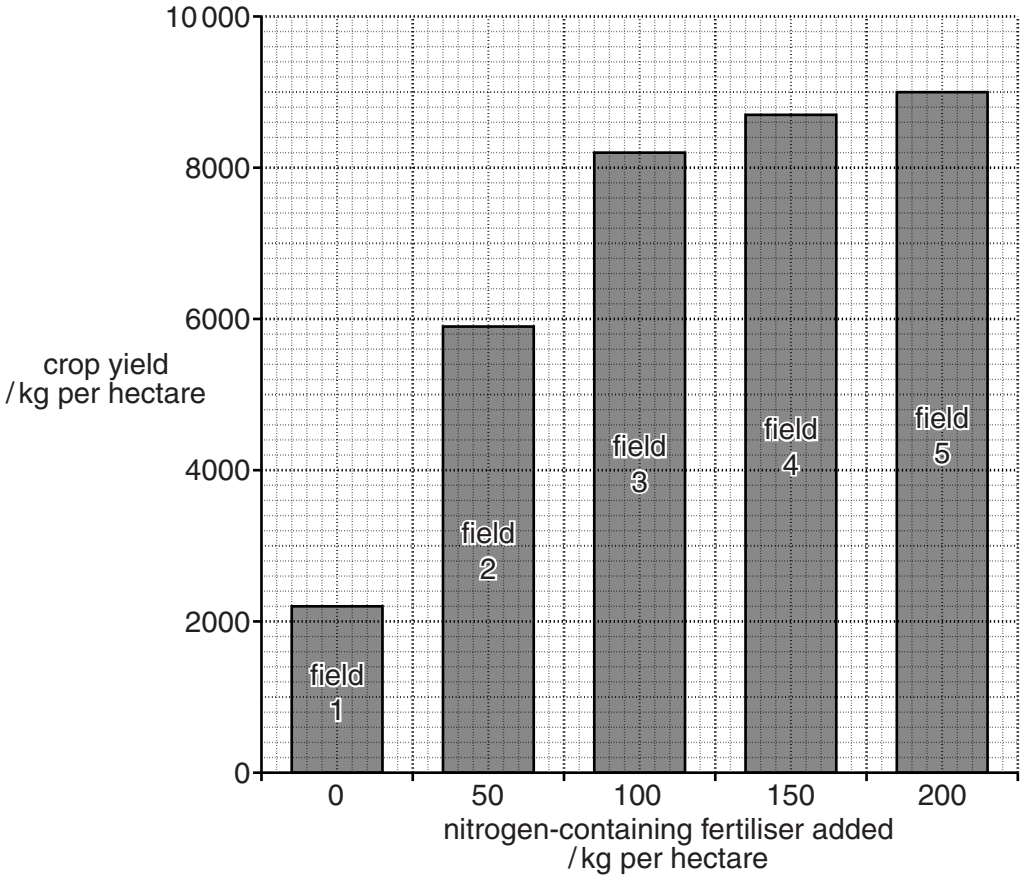


Fig. 13.1

(a) Explain how nitrogen-containing ions are able to pass from soil into a plant.

.....

.....

..... [2]

(b) (i) Using Fig. 13.1, state the crop yield when no nitrogen-containing fertiliser is added.

..... kg per hectare [1]

(ii) Calculate the **increase** in crop yield when 100 kg per hectare of nitrogen-containing fertiliser is used, rather than 50 kg per hectare.

Show your working.

..... kg per hectare [2]

(iii) Explain why the addition of nitrogen-containing fertiliser produces an increase in the yield of maize.

.....

.....

.....

..... [2]

(iv) Use Fig. 13.1 to suggest the crop yield when 250 kg per hectare of nitrogen-containing fertiliser is used.

..... kg per hectare [1]

(c) Explain why most forms of life are dependent on plants carrying out photosynthesis.

.....

.....

.....

..... [2]

[Turn over

14 The first member of the alkene homologous series is ethene.

Ethene is an unsaturated hydrocarbon.

For Examiner's Use

(a) State the general formula of the alkenes.

..... [1]

(b) Ethene reacts with hydrogen to form ethane.

(i) State the type of reaction that takes place when ethene reacts with hydrogen.

..... [1]

(ii) State, in terms of bonds, how the structure of ethene differs from ethane.

..... [1]

(c) Ethene undergoes polymerisation to form poly(ethene).

Draw the structure of poly(ethene).

[2]

15 **(a)** What is coronary heart disease?

..... [1]

(b) State **two** causes of coronary heart disease.

1. .....

.....

2. .....

.....
[2]

16 Fig. 16.1 shows a vernier scale and a micrometer scale.

For Examiner's Use

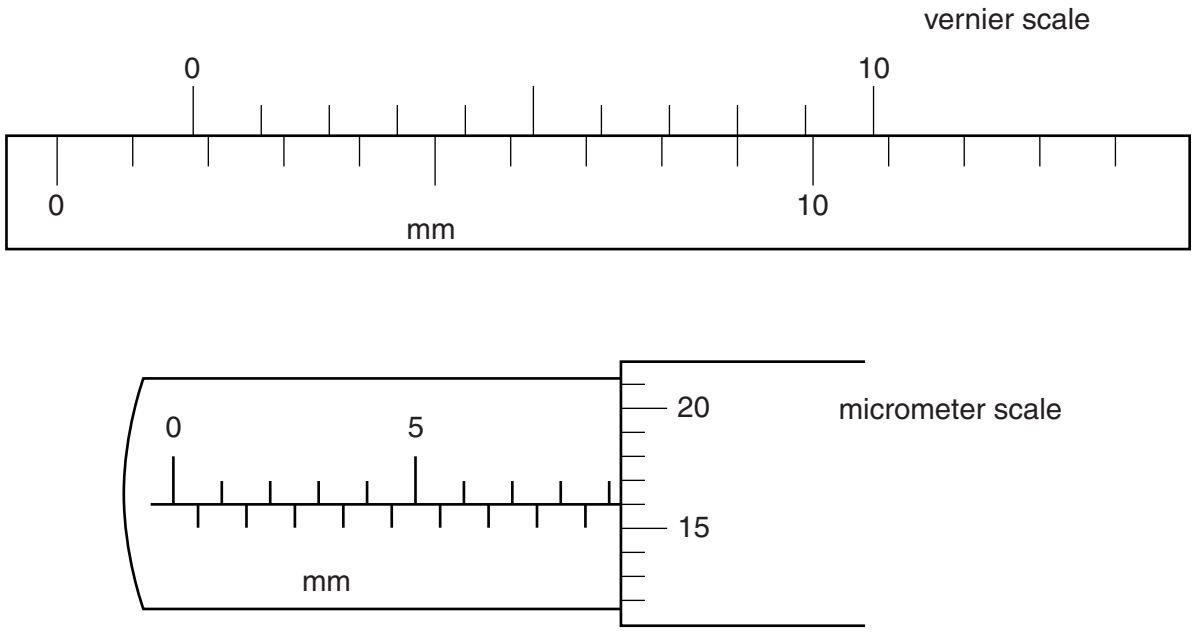


Fig. 16.1

(a) The reading on the vernier scale is .....mm. [1]

(b) The reading on the micrometer scale is .....mm. [1]

17 A bar is placed on a pivot and blocks of mass m_1 and m_2 are placed on the bar, as shown in Fig. 17.1.

For Examiner's Use

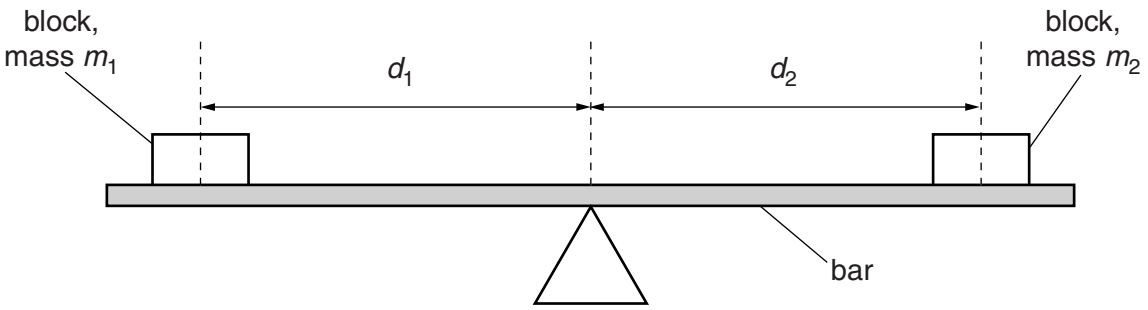


Fig. 17.1

The bar is horizontal.

The distances d_1 and d_2 of the blocks from the pivot are shown in Fig. 17.1.

The masses and their distances from the pivot may be changed so that the bar stays horizontal, tips clockwise or tips anticlockwise.

Fig. 17.2 shows the bar tipping anticlockwise.

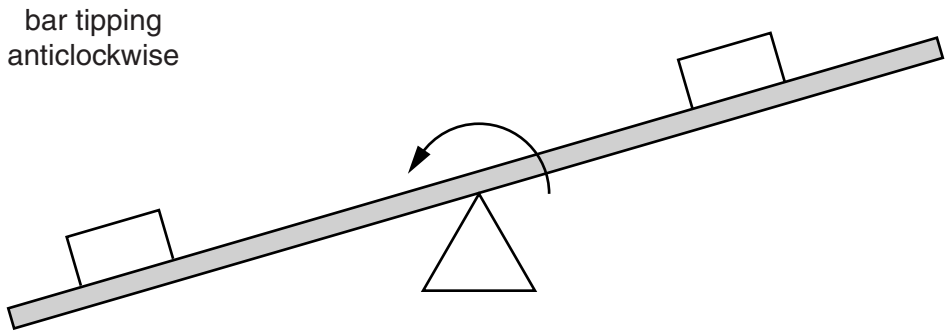


Fig. 17.2

Different masses m_1 and m_2 and distances d_1 and d_2 are shown in Table 17.1.

For Examiner's Use

Complete Table 17.1 by stating whether the bar is horizontal, tips clockwise or tips anticlockwise. The first line has been completed for you.

Table 17.1

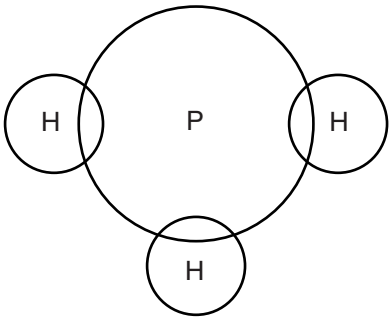
m_1 / g	d_1 / cm	m_2 / g	d_2 / cm	horizontal, tips clockwise, tips anticlockwise.
20	15	20	15	horizontal
20	15	20	20	
30	15	20	15	
10	15	5	15	
30	10	25	12	

[2]

18 (a) Phosphine contains phosphorus and hydrogen and has the formula PH_3.

Phosphorus is in Group V of the Periodic Table.

Complete Fig. 18.1 to show the arrangement of the outer-shell electrons in a molecule of phosphine.



[2]

Fig. 18.1

(b) Complete the following sentences.

The type of bonding present in phosphine is ..... .

Compounds with this type of bonding have ..... melting points and are formed when a ..... combines with a ..... . [3]

19 Fig. 19.1 shows a pin in front of a plane mirror.

A ray of light is incident on the mirror as shown.

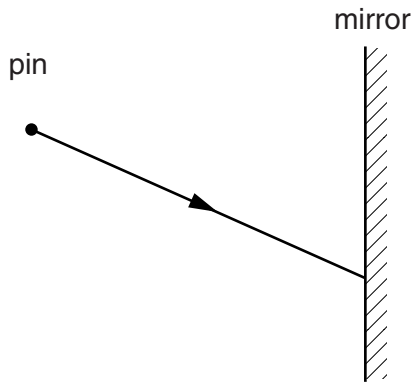


Fig. 19.1

On Fig. 19.1,

(a) draw the normal at the point where the ray is incident on the mirror, [1]

(b) draw the reflected ray, [1]

(c) mark the position of the image of the pin with an **X**. [1]

20 The following is a list of gases.

acetylene **ammonia** **carbon dioxide** **carbon monoxide**

ethane **nitrogen** **oxygen** **sulfur dioxide**

Use the list to complete the following sentences.

Each gas may be used once, more than once or not at all.

(a) The gas produced by complete combustion of hydrocarbon fuels is ..... . [1]

(b) The gases used in welding are ..... and ..... . [1]

(c) The gas that makes up most of the air is ..... . [1]

(d) The gas that turns Universal Indicator red is ..... . [1]

DATA SHEET
The Periodic Table of the Elements

Group																	
I	II										III	IV	V	VI	VII	0
							1 H Hydrogen 1									4 He Helium 2
7 Li Lithium 3	9 Be Beryllium 4										11 B Boron 5	12 C Carbon 6	14 N Nitrogen 7	16 O Oxygen 8	19 F Fluorine 9	20 Ne Neon 10
23 Na Sodium 11	24 Mg Magnesium 12										27 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulfur 16	35.5 Cl Chlorine 17	40 Ar Argon 18
39 K Potassium 19	40 Ca Calcium 20	45 Sc Scandium 21	48 Ti Titanium 22	51 V Vanadium 23	52 Cr Chromium 24	55 Mn Manganese 25	56 Fe Iron 26	59 Co Cobalt 27	59 Ni Nickel 28	64 Cu Copper 29	65 Zn Zinc 30	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36
85 Rb Rubidium 37	88 Sr Strontium 38	89 Y Yttrium 39	91 Zr Zirconium 40	93 Nb Niobium 41	96 Mo Molybdenum 42	Tc Technetium 43	101 Ru Ruthenium 44	103 Rh Rhodium 45	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	128 Te Tellurium 52	127 I Iodine 53	131 Xe Xenon 54
133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57 *	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	186 Re Rhenium 75	190 Os Osmium 76	192 Ir Iridium 77	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	209 Po Polonium 84	210 At Astatine 85	222 Rn Radon 86
223 Fr Francium 87	226 Ra Radium 88	227 Ac Actinium 89 †															

140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	147 Pm Promethium 61	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	159 Tb Terbium 65	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71
232 Th Thorium 90	231 Pa Protactinium 91	238 U Uranium 92	237 Np Neptunium 93	244 Pu Plutonium 94	243 Am Americium 95	247 Cm Curium 96	247 Bk Berkelium 97	251 Cf Californium 98	252 Es Einsteinium 99	257 Fm Fermium 100	258 Md Mendelevium 101	259 No Nobelium 102	260 Lr Lawrencium 103

* 58–71 Lanthanoid series
† 90–103 Actinoid series

Key

a
X
b

a = relative atomic mass
X = atomic symbol
b = atomic (proton) number

The volume of one mole of any gas is 24 dm^3 at room temperature and pressure (r.t.p.).